

(20) **Today**

5.3 Heteronuclear Diatomic Molecules

5.4 Polyatomic Molecules

(22) **Second Class from Today**

Chap 6 Acids and Bases

**Next Class (21)**

5.4 Polyatomic Molecules

Chap 6 Acids and Bases

~~Third Class from Today (23)~~

Test 2 on Chap 4 and 5

*Test 2 postponed  
until Nov. 6.*

Introduce MOs (s, p, d orbital interactions)

Diatomic Molecules and Orbital Mixing

Heteronuclear Diatomic Molecules

Polyatomic molecules

# Interpreting the MO diagram HF

Energy levels are drawn higher on the page for high  $E e^-$  ... and lower for low  $E e^-$

If you don't have the energies you can use electronegativity trends as a standing

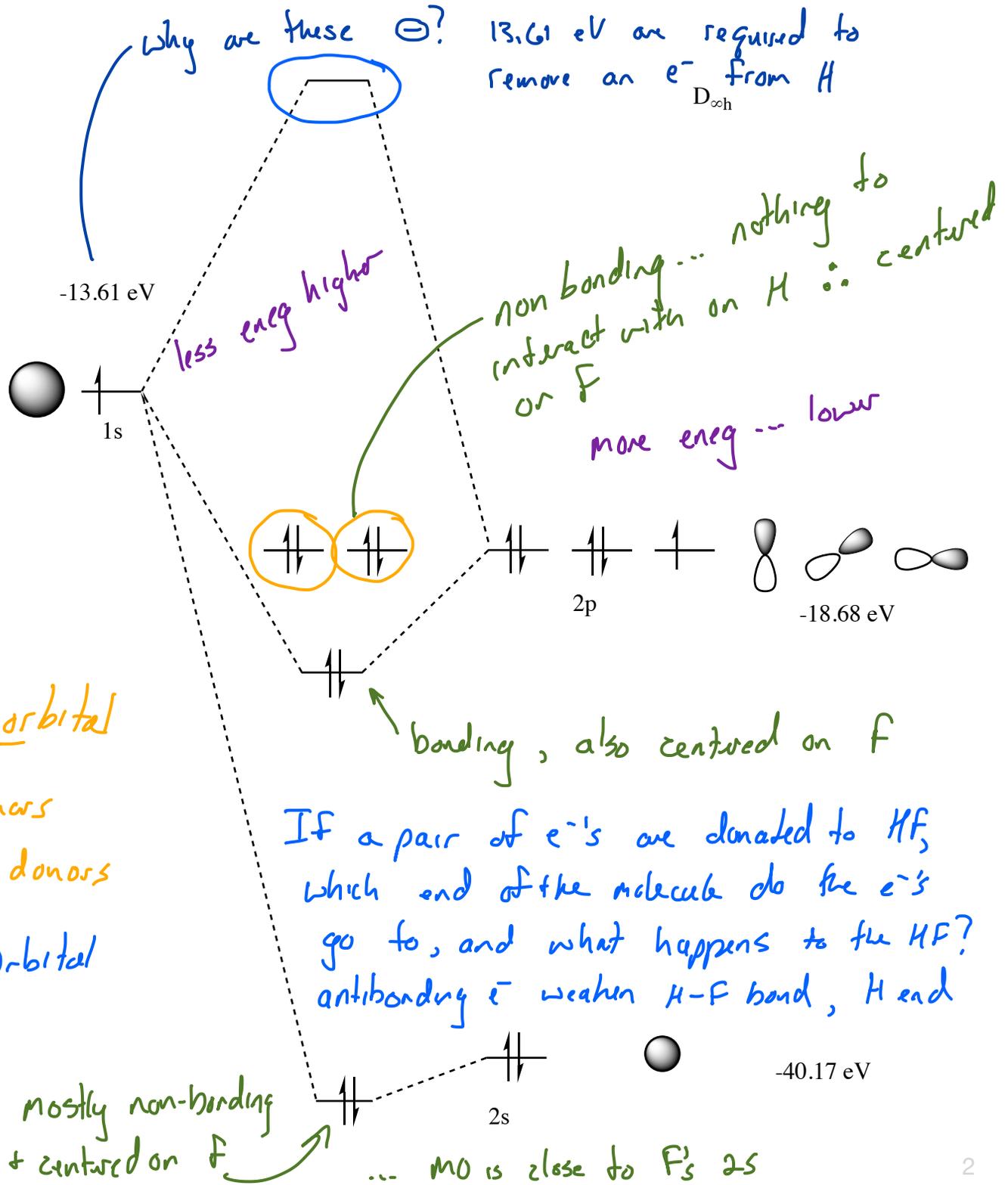
## Interpreting

Highest occupied molecular orbital

HOMO nucleophiles -  $e^-$  donors  
Lewis Bases -  $e^-$  pair donors

Lowest unoccupied Molecular Orbital

LUMO accepts  $e^-$  density  
electrophiles  
Lewis acids



# Interpreting the MO diagram



anti bonding centered on C

O centered bonding orbitals

HOMO is C centered

LUMO is C centered

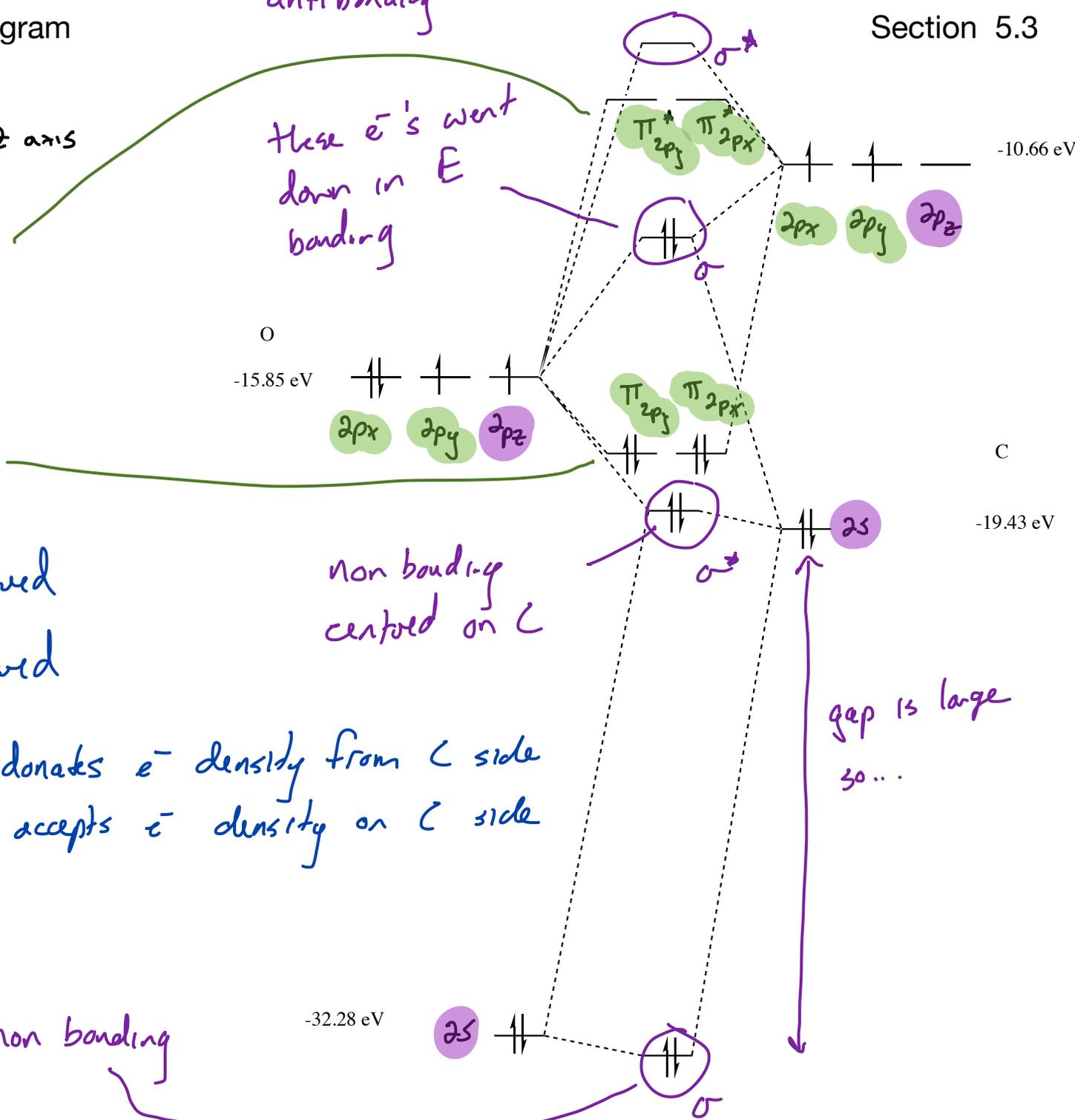
carbon monoxide donates e<sup>-</sup> density from C side  
 carbon monoxide accepts e<sup>-</sup> density on C side

antibonding

these e<sup>-</sup>'s went down in E bonding

non bonding centered on C

O centered mostly non bonding

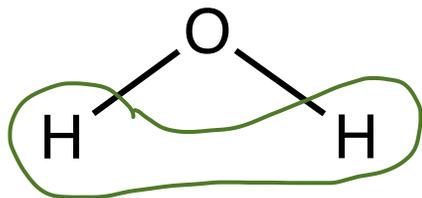


MO Diagram for H<sub>2</sub>O - polyatomic molecules

Section 5.8

central atom

outer atoms



2H

we no longer  
have 2 atoms  
2 sides

2p — — —

— —

group orbitals

"SALC"

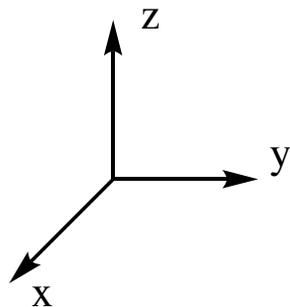
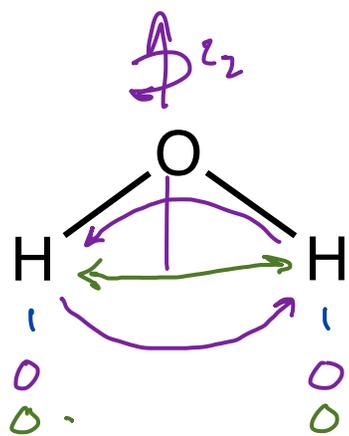
↑  
symmetric

2s —

MO Diagram for H<sub>2</sub>O: SALCs

We use character tables to determine the symmetry of the SALCs formed from H's 1s orbitals

Section 5.3



C <sub>2v</sub>	E	C <sub>2</sub>	σ <sub>v</sub> (xz)	σ <sub>v</sub> (yz)		
A <sub>1</sub>	1	1	1	1	z	x <sup>2</sup> , y <sup>2</sup> , z <sup>2</sup>
A <sub>2</sub>	1	1	-1	-1	R <sub>z</sub>	xy
B <sub>1</sub>	1	-1	1	-1	x, R <sub>y</sub>	xz
B <sub>2</sub>	1	-1	-1	1	y, R <sub>x</sub>	yz

$\Gamma = 2 \text{ (purple)} + 0 \text{ (green)} + 2 \text{ (orange)}$

$\Gamma = A_1 + B_2$

1. Find point group for molecule
2. Find reducible representation for orbitals that make up the SALC
3. Find irreducible representations that combine to form the reducible representation

2 1s's → 2 SALCs with A<sub>1</sub> + B<sub>2</sub>