Report Form

Activity 11: Mass and Chemical Equations

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$$AgNO_3(aq) \ + \ NaCl(aq) \longrightarrow AgCl(s) \ + \ NaNO_3(aq)$$
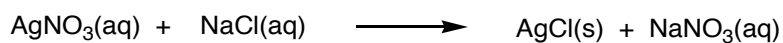
If the equation written above is not balanced, balance it.

Determine the molar mass of $AgNO_3$, $NaCl$, $AgCl$, and $NaNO_3$.

Provided that there is enough $AgNO_3$ present, 1 mol of NaCl will produce how many moles of AgCl.

What is the mass ratio (NaCl:AgCl) for this reaction?

If 1.04 g of AgCl was produced, how much NaCl was originally present?

$$Pb(NO_3)_2(aq) \ + \ NaCl(aq) \ \longrightarrow \ PbCl_2(s) \ + \ NaNO_3(aq)$$

If the equation written above is not balanced, balance it.

Determine the molar mass of all of the reactants and products.

1 mole of $Pb(NO_3)_2$ reacts with how many moles of NaCl?

If 4.05 g of $Pb(NO_3)_2$ must be precipitated from solution, how much NaCl should be added?

salicylic acid   acetic anhydride   acetylsalicylic acid
                                                          aspirin
                                                                    acetic acid

Assuming enough acetic anhydride is added, how many grams of aspirin could be made from 1.50 g of salicylic acid?