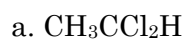
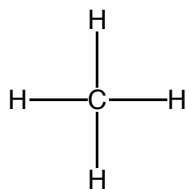


1. Write the ground state electron configuration for N.
2. Draw an energy level diagram for the valence electrons of O.
3. Draw Lewis structures for the following molecules

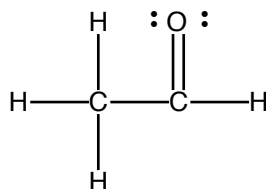


4. Draw three dimensional representations (wedge and dash bond drawings without electron dots) of the following Lewis Structures

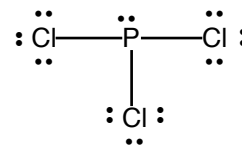
a.



b.

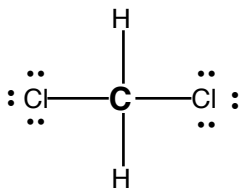


c.

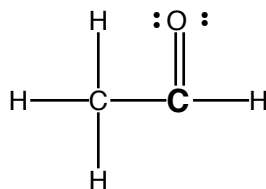


5. Indicate the hybridization of the emboldened atoms in the following Lewis structures

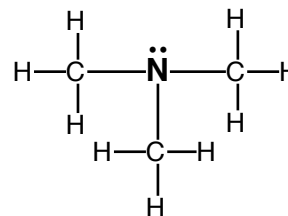
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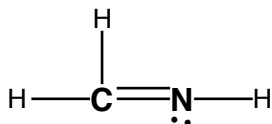
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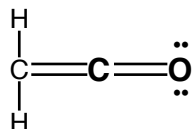
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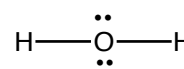
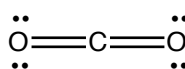
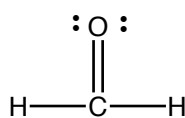
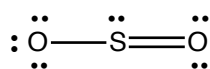
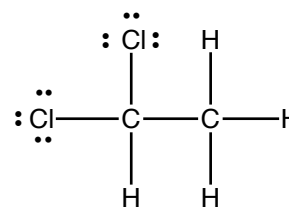
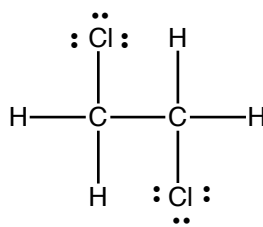
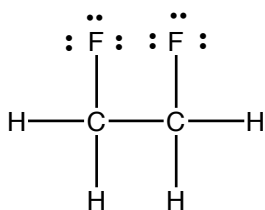
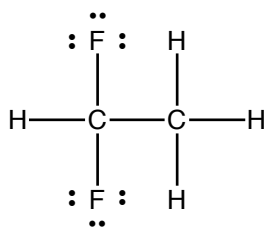
d.



e.

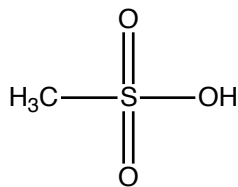
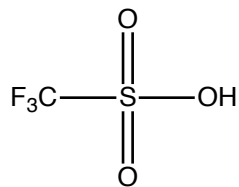


6. Several Lewis structures are drawn below. Circle the polar molecules.

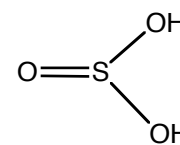
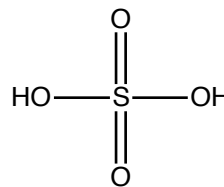


7. For each pair of acids identify the stronger acid

a.



b.



c.



d.

