Next Class (9)

Sections 1.12 Drawing Chemical Structures

Sections 2.1 - 2.4 Polar Covalent Bonds, Formal Charges, Resonance/Electron Delocalization

Bring Modeling Kits

Third Class from Today (11)

Sections 2.4 – 2.6 Resonance/Electron Delocalization

> Sections 2.7 – 2.11 Acids and Bases

(10) Second Class from Today

Sections 2.4 – 2.6 Resonance/Electron Delocalization

Bring Modeling Kits

(8) **Today**

Sections 1.12 Drawing Chemical Structures Lewis & Kekulé Structures

Section 1.12

H H H H:C:O:C:C:H H H H

Chemists use different drawings to place emphasis on different aspects of a molecule.

Representations are used to solve typographical issues.

Molecular Formulas as Compared to Condensed Structures/Structural Section 1.12 Formulas

 C_3H_8O

Molecular Formulas as Compared to Condensed Structures/Structural Section 1.12 Formulas

In organic, condensed structures typically start with a C, and everything immediately to the right of the C is connected to that first C. When the the first C is finally connected to the second C, now that atoms right of the second C are connected to second C. In acyclic unbranched molecules atoms to the right of the second C are not connected to the first C.

 C_3H_8O $CH_3OCH_2CH_3$ $CH_3CH_2OCH_3$ $CH_3CH_2CH_2OH$ $CH_3CHOHCH_3$

In organic, condensed structures typically start with a C, and everything immediately to the right of the C is connected to that first C. When the the first C is finally connected to the second C, now that atoms right of the second C are connected to second C. In acyclic unbranched molecules atoms to the right of the second C are not connected to the first C.

CH₂CHCH₃

Because bonds are not drawn, condensed structures require the reader to bring some chemical knowledge to their interpretation.

Section 1.4



 $CH_3CH(OH)CH_2CH_3$ $CH_3(CH_2)_3CH_3$ $CH_3CH_2CH(CH_3)_2$

Parentheses () in structures are typically used to set off side chains, to indicate a repeating unit, or to indicate multiple groups of the same structure.

Condensed Structures/Structural Formulas

Often, chemists omit parentheses when they are not absolutely necessary,

ц /		L \	\mathbf{C}	ப
⊓3(ĹΟΓ	₂ 2)	30	пз

CH₃CHOHCH₃ CH₃CH(OH)CH₃

 $CH_3COCH_2CH_3$ $CH_3C(O)CH_2CH_3$

and sometimes chemists do things for aesthetic reasons.

 $C(CH_3)_3OH$ $CH_3C(CH_3)_2OH$ $(CH_3)_3COH$

CH₃CHOHCH₂CH₃

 $CH_3C(O)CH(CH_3)_2$

CH₃CHO

When a bond ends and the atom isn't labeled it is assumed to be C.

When there aren't enough bonds drawn to a C atom, the "missing" bonds are C atom to H atom bonds.

All other atoms are labeled.

Heptane

2-heptanol

Different structures serve different purposes, but they represent the same things

Converting Between Structure Types





$CH_{3}CH(OH)CH_{2}CH(CH_{3})CH_{2}CH_{3}$