

Name \_\_\_\_\_

Test 4

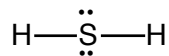
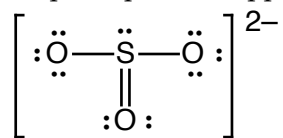
1. Draw Lewis structures for the following compounds (include all resonance structures and formal charges even if the formal charge is 0).

a.  $\text{BrO}_2^-$

b.  $\text{F}_3\text{PS}$

c.  $\text{OCS}$

2. Draw wedge and dashed bond shapes for the following molecules. There are two criteria for grading this problem: 1. is the correct shape drawn, and 2. the quality of the drawing. If necessary name the shape or provide approximate bond angles.



7. (2 pts. ea.) Indicate whether the following molecules are polar (Lewis structures are provided)

